Cosmo School

Chemistry Revision Worksheet for monthly assessment

Class: 9

MCQ worksheet
Chapter-4 (Periodic table)

1. What is the group number of Mg?
   a) 2
   b) 3
   c) 4
   d) 1
2. What is the position of atom with 17 electrons?
   a) 2^{nd} period, group-2
   b) 2^{nd} period, group-7
   c) 3^{rd} period, group-5
   d) 3^{rd} period, group-7
3. Which one is of the same period as Bromine?
   a) 2,5
   b) 2,8,7
   c) 2,7
   d) 2,8,8,2
4. Which level does the last electron of Cl enters into?
   a) 1s
   b) 3p
   c) 2s
   d) 3d
5. Which elements valency is zero?
   a) Na
   b) Ni
   c) Ne
   d) Fe
6. Which of the two elements gain the electronic configuration of argon to form ion?
   a) Ca, Cl
   b) Sc, Cl
   c) K, O
   d) S, C
7. How many elements are recognized by IUPAC?
   a) 98
   b) 118
   c) 113
   d) 120
8. Which one is the element of 4^{th} period?
   a) Mg
   b) N
   c) Na
   d) K
9. Which element has the highest ionization energy?
   a) S
   b) Si
   c) P
   d) Al
10. Which physical condition is required for ionization energy?
    a) Solid
    b) Gaseous
    c) Liquid
    d) Plasma
11. Electronegativity of which one is maximum?
    a) O
    b) S
    c) N
    d) F
12. Which reactivity series is correct?
    a) Na>Mg>Al
    b) Al>Mg>Na
    c) K<Na<Li
    d) Mg>Ca>Sr
13. Which of the following is metaloid?
    a) S
    b) Si

c) O
d) Mg

14. Hydrogen is located in group-1, because—
    i) It forms halide like metal
    ii) It is electropositive
    iii) It forms hydride with metal
    Which one is correct?
    a) i & ii
    b) ii & iii
    c) i & iii
    d) i, ii &iii

15. What is the main basis of modern periodic table?
    a) Electronic configuration
    b) Atomic number
    c) Atomic mass
    d) Relative atomic mass

Chapter- 6 (Concept of mole and chemical counting)

1. What is the molar volume of 16g oxygen at STP?
    a) 5.6 L
    b) 22.4 L
    c) 11.2 L
    d) 33.6 L
2. If 0.1 mole solute get dissolved in one liter solution, then what will be the concentration of the solution?
    a) 0.1 M
    b) 0.01 M
    c) 1 M
    d) 0.5 M
3. What is the concentration in molarity of decimolar solution?
    a) 0.7 molar
    b) 0.1 molar
    c) 0.5 M
    d) 0.01 M
4. If water is used as a solvent the solution produced is called—
    a) Aqueous solution
    b) Alkaline solution
    c) Dilute solution
    d) Concentrated solution
5. 1 mole H atom equals to –
    i) 1.008g H atoms
    ii) 6.02×10^{23} H atoms
    iii) 22.4 L H atoms
    Which one is correct?
    a) i & ii
    b) ii & iii
    c) i & iii
    d) i, ii &iii
6. How many liters of solution will be produced from 100g limestone with molarity 0.5 M?
    a) 1 L
    b) 4 L
    c) 2 L
    d) 10 L
7. What is the volume of 44g CO_2 and 32g O_2?
    a) Volume of CO_2 > volume of O_2
    b) Volume of CO_2 < volume of O_2
    c) Volume of CO_2 = volume of O_2
    d) Volume of $CO_2 \neq$ volume of O_2
8. To form CO_2 molecule, how much oxygen will react with 3g carbon?

a) 8g
b) 32g
c) 12g
d) 44g

9. How many ions are present in 1 mole of Na^+ ion?
a) $2 \times 6.02 \times 10^{23}$
b) 3.1416
c) 6.02×10^{23}
d) 6.623×10^{11}

Read the following stem and answer the question no. 10 and 11:
At STP volume of 10g of X gas is 112L.

10. Which one is the X gas?
a) H_2
b) O_2
c) CO_2
d) NH_3

11. The gas X —
i) Molecular mass is 2
ii) At STP molar volume is 22.4 L
iii) Atomic mass is 2
Which one is correct?
a) i & ii
b) ii & iii
c) i & iii
d) i, ii &iii

Read the following stem and answer the question no. 12 and 13:
In a 200 mL container 10.6g X is taken and water is added. When the volume of the solution is 200 mL, semimolar solution is produced.

12. What is the compound X?
a) NaOH
b) $Mg(OH)_2$
c) Na_2CO_3
d) $Ca(OH)_2$

13. What will be the volume of the solution if 20g X dissolves in 0.75M solution?
a) 200 mL
b) 500 mL
c) 100 mL
d) 250 mL

14. How many molecules are present in 24.5g H_2SO_4?
a) 12.5×10^{22}
b) 1.15×10^{23}
c) 1.505×10^{23}
d) 1.198×10^{23}

15. The mass of 3.01×10^{23} atoms of carbon is --
a) 6g
b) 6.22g
c) 12g
d) 12.22g